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QUIZZES

Practice Test-1(s and p Block Elements)



10 Questions



7 min

Topics

Periodic Trends in Physical Properties

Start Quiz

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06 : 57



1/10



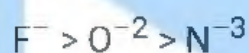
7 min



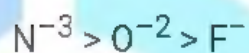
Hint

Q : The correct order of ionic radius is

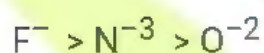
A



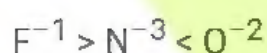
B



C



D



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1

2

3

4

5

6

7

06 : 53



2/10



7 min



Hint

Q : Which is an incorrect order of atomic radii

A

Mg > Sr

B

Be > C

C

Al > S

D

K > Ar

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1

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06 : 50



3/10



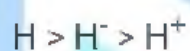
7 min



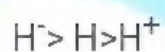
Hint

Q : Which of the following order of ionic radii is correct

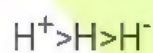
A



B



C



D



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1

2

3

4

5

6

7

06 : 47



4/10



7 min



Hint

Q : Which atom has the smallest size



C



F



He



N

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1

2

3

4

5

6

7

06 : 43



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5/10



7 min



Hint

Q : For which molecule the bonding pair of electrons is equally shared between the atoms?



HF



HCl



H₂



BF₃

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1

2

3

4

5

6

7

06 : 39



6/10



7 min



Hint

Q : An element with highest first ionization energy

A

Nitrogen

B

Boron

C

Oxygen

D

Beryllium

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1

2

3

4

5

6

7

06 : 36



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7/10



7 min



Hint

Q : Which factor does not effect ionization energy across the period

A

Number of shell

B

Nuclear Charge

C

Nature of orbital

D

Both A & C

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1

2

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4

5

6

7

06 : 31



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8/10



7 min



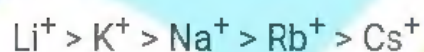
Hint

Q : Which of following is a correct order of degree of hydration in alkali metal ions

A



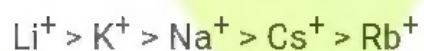
B



C



D



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4

5

6

7

8

9

10

06 : 29



9/10



7 min



Hint

Q : Select the correct increasing order of melting point of alkaline earth metals

A

Be > Mg > Ca > Sr

B

Be > Ca > Sr > Mg

C

Mg > Be > Ca > Sr

D

Mg > Sr > Ca > Be

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4

5

6

7

8

9

10

Q : Which of the following is shows maximum electrical conductance:

- ☐ Na
- ☐ Al
- ☐ Mg
- ☐ All of these

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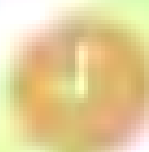
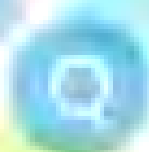
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QUIZ RESULT

Practice Test-1(s and p Block Elements)



Time



Score



0/10

0%

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Correct



Unattempted

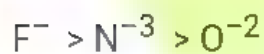
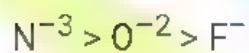
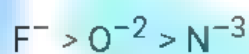


Incorrect



1/10

Q : The correct order of ionic radius is



Explanation

Along the period anionic radius decreases from left to right



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correct



2/10

Q : Which is an incorrect order of atomic radii



Mg > Sr



Be > C



Al > S



K > Ar

Explanation

Atomic radii decrease from left to right and increase from top to bottom.



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Correct



Question



Incorrect



3/10

Q : Which of the following order of ionic radii is correct



$H > H^- > H^+$



$H > H > H^+$



$H^+ > H > H^-$



$H^- > H^+ > H$

Explanation

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Anion is greater and cation smaller than its parent atom.



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correct



4/10

Q : Which atom has the smallest size



C



F



He



N

Explanation

Less no. of shells in He than other



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Correct



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Incorrect



5/10

Q : For which molecule the bonding pair of electrons is equally shared between the atoms?



HF



HCl



H₂



BF₃

Explanation

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In H₂ molecule both atoms have same electronegativity

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correct



6/10

Q : An element with highest first ionization energy



Nitrogen



Boron



Oxygen



Beryllium

Explanation

Along the period IE increases from left to right but there are abnormal trends

$V A > VI A, II A > III A$

Due to stable electronic configuration



Correct



Explanation



Incorrect



7/10

Q : Which factor does not effect ionization energy across the period



Number of shell



Nuclear Charge



Nature of orbital



Both A & C

Explanation

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In period number of shell and shielding effect will remain constant



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Correct



Unattempted



Incorrect



8/10

Q : Which of following is a correct order of degree of hydration in alkali metal ions



$\text{Li}^+ > \text{Na}^+ > \text{K}^+ > \text{Rb}^+ > \text{Cs}^+$



$\text{Li}^+ > \text{K}^+ > \text{Na}^+ > \text{Rb}^+ > \text{Cs}^+$



$\text{Li}^+ < \text{Na}^+ < \text{K}^+ > \text{Rb}^+ < \text{Cs}^+$



$\text{Li}^+ > \text{K}^+ > \text{Na}^+ > \text{Cs}^+ > \text{Rb}^+$

Explanation

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as size increases down the group degree of hydration decreases.



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Correct



Unattempted



Incorrect



9/10

Q : Select the correct increasing order of melting point of alkaline earth metals



Be > Mg > Ca > Sr



Be > Ca > Sr > Mg



Mg > Be > Ca > Sr



Mg > Sr > Ca > Be

Explanation

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In alkaline earth melting point decrease top to bottom due to increasing size but Mg show abnormal trend due close hexagonal packing



correct



10/10

Q : Which of the following is shows maximum electrical conductance:



Na



Al



Mg



All of these

Explanation

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Electrical conductance increases from IA to IIIA.



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QUIZZES

Practice Test-2(s and p Block Elements)

10 Questions

75 Min

10 Marks

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Q : Electronic configuration of alkali metals is_____

☐ ns^2, np^3

☐ ns^2, np^1

☐ ns^2, np^0

☐ ns^1

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SAEED MDCAT TEAM

f SAEEDMDCAT

Q : One of the following is least reactive among following

☐ Li

☐ Na

☐ K

☐ Cs

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SAEED MDCAT TEAM

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Q : One of the following will form superoxide when burnt in air

☐ Li

☒ Na

☐ K

☐ Be

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SAEED MDCAT TEAM

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Q : Which one of the following is most basic oxide

☐ BeO

☐ MgO

☐ SrO

☐ CaO

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SAEED MDCAT TEAM

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Q : Which IIA hydroxide decomposes most easily

☐ $\text{Mg}(\text{OH})_2$

☐ $\text{Ca}(\text{OH})_2$

☐ $\text{Sr}(\text{OH})_2$

☐ $\text{Ba}(\text{OH})_2$

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Q : Hydroxide of which Alkaline earth metal cannot be prepared by direct reaction with water.

☐ Be

☒ Ca

☐ Sr

☐ Ba

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Q : $M + O_2 \rightarrow MO_2$.The metal M is

☐ Be

☐ Mg

☐ Ca

☐ Ba

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Q : Burning of Be in air requires

- ☐ 600°C
- ☐ 800°C
- ☐ 1400°C
- ☐ 200°C

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SAEED MDCAT TEAM

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Q : Which one of the following from 3rd period doesn't react with water



Be



Ca



Na



Ar

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SAEED MDCAT TEAM



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Q : Which of the following element of 3rd period and group III does not react with water



B



Mg



Al



In

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SAEED MDCAT TEAM



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QUIZ RESULT

Practice Test-2(s and p Block Elements)



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correct



1/10

Q : Electronic configuration of alkali metals is_____



ns^2, np^3



ns^2, np^1



ns^2, np^0



ns^1

Explanation

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Electronic configuration of alkali metals is ns^1 .



SAEEDMDCAT



correct



2/10

Q : One of the following is least reactive among following



Li



Na



K



Cs

Explanation

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Due to high I.E Li is least reactive.



SAEEDMDCAT



Correct



Explanation



Incorrect



3/10

Q : One of the following will form superoxide when burnt in air



Li



Na



K



Be

Explanation

Potassium will form superoxide among all given options, due to high reactivity.



correct



4/10

Q : Which one of the following is most basic oxide



BeO



MgO



SrO



CaO

Explanation

In group IA & IIA basic character of their oxides increases down the group.



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correct



5/10

Q : Which IIA hydroxide decomposes most easily



Mg(OH)_2



Ca(OH)_2



Sr(OH)_2



Ba(OH)_2

Explanation

Ease of decomposition is inversely related to cationic size inc. order of cationic size



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correct



6/10

Q : Hydroxide of which Alkaline earth metal cannot be prepared by direct reaction with water



Be



Ca



Sr



Ba

Explanation

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Be does not react with water



SAEEDMDCAT



Practice Test-2(s and p Block Elements)



Correct



Unattempted



Incorrect



7/10

Q : $M + O_2 \rightarrow MO_2$.The metal M is

A

Be

B

Mg

C

Ca

D

Ba

Explanation

MO_2 = Peroxide Ba gives peroxide in air at $500^\circ - 600^\circ C$



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1

2

3

4

5

6

7



Practice Test-2(s and p Block Elements)



Correct



Unattempted



Incorrect



8/10

Q : Burning of Be in air requires

A

600°C

B

800°C

C

1400°C

D

200°C

Explanation

Be is the least reactive member of IIA and burns at 800°C





Practice Test-2(s and p Block Elements)



Correct



Unattempted



Incorrect



9/10

Q : Which one of the following from 3rd period doesn't react with water

A

Be

B

Ca

C

Na

D

Ar

Explanation

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Argon does not react directly with water.



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Practice Test-2(s and p Block Elements)



Correct



Unattempted



Incorrect



10/10

Q : Which of the following element of 3rd period and group III does not react with water

A

B

B

Mg

C

Al

D

In

Explanation

Al does not react with water, its oxides are amphoteric